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Common
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and formula of some
chemical compounds..

Valence Electrons and
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Terms. annearnold.

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leahrosner. HD 1150

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leahrosner. Specialized
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STUDY. PLAY.

Monatomic Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

~~Chapter 9: Chemical~~

Page 12/35

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~~Names and Formulas~~

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To name a binary molecular compound, use the following guidelines: 1. Write the names of the elements in the order listed in the formula 2. Use prefixes appropriately to indicate the number of each kind of atom 3. End the name of the second element with the suffix -ide

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Chemical

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First, confirm that the

compound is a binary
molecular compound.

Name the elements in the
order listed in the

formula. Use prefixes to
indicate the number of

each kind of atom. Omit
the prefix -mono when

the formula contains only

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one atom of the first element in the name. The suffix of the name of the second element is -ide.

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1. when the name of the anion (X) ends in -ide the acid name begins with the prefix hydro-.

The stem of the anion has the suffix -ic and is

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followed by the word acid (ex. HCl = hydrochloric acid) 2. when the anion name ends in -ite the acid name is the stem of the anion with the suffix -ous followed by the word acid (ex. H_2SO_3 -- X =sulfite = sulfurous acid)

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~~Test Review ...~~

Writing the Names and Formulas for Bases. A base is a compound that produces OH^- in water.

When naming a base, you name it like any other compound that starts with a regular or transition metal. Ex:

NaOH = sodium hydroxide. When writing the formula for a base, remember to balance

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charges. Ex: magnesium
hydroxide = $\text{Mg}(\text{OH})_2$

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SECTION 9.3 NAMING
AND WRITING
FORMULAS FOR
MOLECULAR

COMPOUNDS (pages
268 – 270) This section
explains the rules for
naming and writing
formulas for binary

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molecular compounds.

Naming Binary

Molecular Compounds

(pages 268 – 269) 1.

Circle the letter of the
type(s) of elements that
form binary molecular
compounds.

~~Name Date Class~~

~~CHEMICAL NAMES~~

~~AND FORMULAS 9~~

~~SECTION 9.2 NAMING~~

~~AND WRITING~~

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FORMULAS FOR IONIC COMPOUNDS

1. Write the formulas for these binary ionic

compounds. c.

potassium iodide 2. write

the formulas for the

compounds formed c. b.

f. 3. Name the following

binary ionic compounds.

sodium sulfide g. CuCl :

h. SnCl_2 . i. a. MnO_2 c. d.

SrBr_2 221 Chapter 9

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CHAPTER 9 Assessment

42. a. 2 – b. 1+ c. 1 – d.

3+ 43. a. 2+ b. 2+ c. 3+

d. 1+ 44. a. barium ion b.

iodide ion c. silver ion d.

mercury(II) ion 45.

cyanide, CN^- and

hydroxide, OH^- 46. a.

hydroxide ion b.

lead(IV) ion c. sulfate ion

d. oxide ion 47. zero; A

compound is electrically

neutral. 48. The symbols

for the cation and

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Monatomic Ion. Consists
of a single atom with a

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positive or negative
charge resulting from the
loss of gain of one or
more valence

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Answer Key~~

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account; E-mail to a
friend; Find other
activities; Start over;

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Help; Use these activities to help yourself to learn the vocabulary and major topics presented in this chapter. A B; anion: any atom or group of atoms that has a negative charge:

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Answer Section

MATCHING 1. ANS: A

PTS: 1 DIF: L1 REF: p.

253 OBJ: 9.1.1 Identify
the charges of

monatomic ions by using
the periodic table, and
name the ions. STA:

Ch.5.a 2.ANS: H PTS: 1

DIF: L1 REF: p. 254

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